PCT WORLD INTELLECTUAL PROPERTY ORGANIZATION
International Bureau

INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification 6 : C07D 401/12 | A1

(11) International Publication Number: **WO 97/41114**

(43) International Publication Date: 6 November 1997 (06.11.97)

(21) International Application Number: PCT/SE97/00674

(22) International Filing Date: 22 April 1997 (22.04.97)

(30) Priority Data:
9601598-7 26 April 1996 (26.04.96) SE

(71) Applicant *(for all designated States except US)*: ASTRA AKTIEBOLAG [SE/SE]; S-151 85 Södertälje (SE).

(72) Inventors; and

(75) Inventors/Applicants *(for US only)*: HÖGBERG, Jan-Åke [SE/SE]; Kämpevägen 41, S-151 54 Södertälje (SE). IOANNIDIS, Panagiotis [GR/SE]; Ovanbygränd 16, S-163 70 Spånga (SE). MATTSON, Anders [SE/SE]; Kopparvägen 188, S-183 46 Täby (SE).

(74) Agent: ASTRA AKTIEBOLAG; Patent Dept., S-151 85 Södertälje (SE).

(81) Designated States: AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, HU, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ARIPO patent (GH, KE, LS, MW, SD, SZ, UG), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG).

Published
With international search report.

(54) Title: PROCESS FOR THE PREPARATION OF A MAGNESIUM SALT OF A SUBSTITUTED SULPHINYL HETEROCYCLE

(57) Abstract

A novel process for the preparation of a magnesium salt of formula (I) of a substituted sulfinyl heterocyclic compound containing an imidazole moiety. The process is carried out by mixing the substituted heterocycle of formula (I) with a weak and a magnesium source. The base and the magnesium source are selected to result in residues which are easy to remove during the reaction. The invention also relates to the use of the produced compounds in medicine.

FOR THE PURPOSES OF INFORMATION ONLY

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AL	Albania	**ES**	Spain	**LS**	Lesotho	**SI**	Slovenia
AM	Armenia	**FI**	Finland	**LT**	Lithuania	**SK**	Slovakia
AT	Austria	**FR**	France	**LU**	Luxembourg	**SN**	Senegal
AU	Australia	**GA**	Gabon	**LV**	Latvia	**SZ**	Swaziland
AZ	Azerbaijan	**GB**	United Kingdom	**MC**	Monaco	**TD**	Chad
BA	Bosnia and Herzegovina	**GE**	Georgia	**MD**	Republic of Moldova	**TG**	Togo
BB	Barbados	**GH**	Ghana	**MG**	Madagascar	**TJ**	Tajikistan
BE	Belgium	**GN**	Guinea	**MK**	The former Yugoslav Republic of Macedonia	**TM**	Turkmenistan
BF	Burkina Faso	**GR**	Greece			**TR**	Turkey
BG	Bulgaria	**HU**	Hungary	**ML**	Mali	**TT**	Trinidad and Tobago
BJ	Benin	**IE**	Ireland	**MN**	Mongolia	**UA**	Ukraine
BR	Brazil	**IL**	Israel	**MR**	Mauritania	**UG**	Uganda
BY	Belarus	**IS**	Iceland	**MW**	Malawi	**US**	United States of America
CA	Canada	**IT**	Italy	**MX**	Mexico	**UZ**	Uzbekistan
CF	Central African Republic	**JP**	Japan	**NE**	Niger	**VN**	Viet Nam
CG	Congo	**KE**	Kenya	**NL**	Netherlands	**YU**	Yugoslavia
CH	Switzerland	**KG**	Kyrgyzstan	**NO**	Norway	**ZW**	Zimbabwe
CI	Côte d'Ivoire	**KP**	Democratic People's Republic of Korea	**NZ**	New Zealand		
CM	Cameroon			**PL**	Poland		
CN	China	**KR**	Republic of Korea	**PT**	Portugal		
CU	Cuba	**KZ**	Kazakstan	**RO**	Romania		
CZ	Czech Republic	**LC**	Saint Lucia	**RU**	Russian Federation		
DE	Germany	**LI**	Liechtenstein	**SD**	Sudan		
DK	Denmark	**LK**	Sri Lanka	**SE**	Sweden		
EE	Estonia	**LR**	Liberia	**SG**	Singapore		

PROCESS FOR THE PREPARATION OF A MAGNESIUM SALT OF A SUBSTITUTED SULPHINYL HETEROCYCLE

Field of the invention.

The present invention relates to a novel process for the preparation of magnesium salts of substituted sulfinyl heterocyclic compounds containing an imidazole moiety as well as the use of the produced magnesium salts in medicine. More particularly, the present invention relates to the preparation of magnesium salts of substituted benzimidazoles such as the magnesium salts of omeprazole and of its single enantiomers.

Background of the invention and prior art.

Substituted benzimidazoles such as for instance the compounds with the generic names omeprazole, lansoprazole, pantoprazole, pariprazole and leminoprazole have properties making the compounds useful as inhibitors of gastric acid secretion. This class of compounds is known as proton pump inhibitors or H^+K^+ATPase inhibitors. There are a large number of patents and patent applications disclosing such proton pump inhibitors and processes for their preparation.

There is a general need in industry that pharmaceutically active compounds should be produced by processes giving products with properties making them suitable for pharmaceutical preparations, such as being easy to handle in a full scale production and having good storage stability.

WO 95/ 01977 discloses a novel magnesium salt of omeprazole with a specific degree of crystallinity making the product suitable for pharmaceutical formulations. The novel product is prepared by a process comprising the following steps; reacting omeprazole with magnesium alcoholate; separating inorganic salts from the reaction mixture; crystallizing the magnesium salt of omeprazole and isolating the product. The magnesium alcoholate is

formed from metallic magnesium which requires special process conditions. The use of magnesium alcoholate in the process constitutes a potential difficulty with the formation of relatively insoluble magnesium salts, such as magnesium hydroxide. Filtration of such magnesium hydroxide is complicated because of gelling and extremely small particle size . The prior process is rather complicated, is water sensitive and requires special conditions. The prior process also has a large equipment requirement in the form of three reaction vessels and a separator. Therefore, there is a need for a more efficient process resulting in shorter manufacturing time, less reaction equipment and giving a higher yield pro volume.

The present invention provides improvements over the process disclosed in WO 95/01977 for the preparation of the magnesium salts of omeprazole and of other substituted benzimidazoles. Process for the preparation of certain salts of the single enantiomers of omeprazole, such as the magnesium salts, and processes for their preparation are described in EP 94917244.9.

As discussed in WO 95/ 01783 the magnesium salts of proton pump inhibitors, such as the magnesium salt of omeprazole, are especially suitable for the manufacturing of pharmaceutical formulations, such as tablets. The magnesium salts are stable, they may be easily purified by crystallization, and are easy to handle in pharmaceutical procedures and processes.

Summary of the invention.

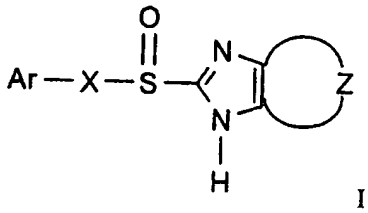
The present invention provides a novel process for the preparation of magnesium salts of substituted sulfinyl heterocycles containing an imidazole moiety and especially of substituted benzimidazole derivatives. The process results in a high yield pro volume, requires less equipment, is less time consuming, environmental friendly and more economically efficient than processes described in the above mentioned patent applications. According to the novel process a magnesium salt of a substituted sulfinyl heterocycle containing an imidazole moiety is prepared by mixing the substituted sulfinyl

heterocycle containing an imidazole moiety with a weak base, preferably an amine or ammonia, and a magnesium source, such as an organic or inorganic magnesium salt or a combination of such salts. By the novel process of the present invention formation of magnesium hydroxide is avoided, for example in the preparation of omeprazole magnesium salt.

Alternatively, the process may also be used to prepare other salts of a substituted sulfinyl heterocycle containing an imidazole moiety, for instance multiple valent salts, such as calcium salts.

Detailed description of the invention.

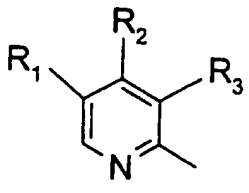
The present invention provides a novel method of preparing a magnesium salt of a substituted sulfinyl heterocycle containing an imidazole moiety with the following formula I.



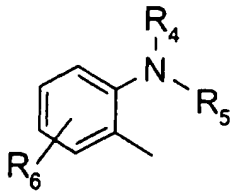
I

wherein

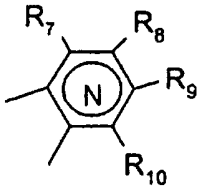
Ar is



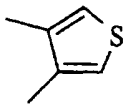
or



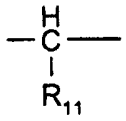
Z is



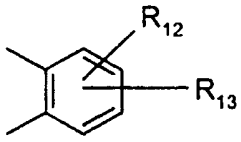
or



and X is



or



wherein

N inside the benzene ring of the benzimidazole moiety means that one of the carbon atoms substituted by R_7-R_{10} optionally may be exchanged for a nitrogen atom without any substituents;

R_1, R_2 and R_3 are the same or different and selected from hydrogen, alkyl, alkylthio, alkoxy optionally substituted by fluorine, alkoxyalkoxy, dialkylamino, piperidino, morpholino, halogen, phenylalkyl and phenylalkoxy; wherein alkyl and alkoxy groups may be branched or linear and may comprise cyclic alkyl groups such as cykloalkylalkoxi groups.

R_4 and R_5 are the same or different and selected from hydrogen, alkyl and aralkyl;

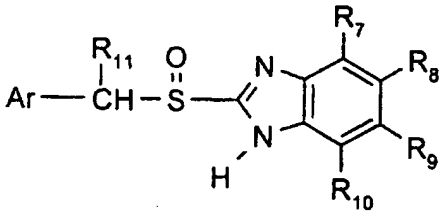
R_6 is hydrogen, halogen, trifluoromethyl, alkyl and alkoxy;

R_7 - R_{10} are the same or different and selected from hydrogen, alkyl, alkoxy, halogen, haloalkoxy, alkylcarbonyl, alkoxycarbonyl, oxazolyl, trifluoroalkyl, or adjacent groups R_7-R_{10} form ring structures which may be further substituted;

R_{11} is hydrogen or forms an alkylene chain together with R_3 and

R_{12} and R_{13} are the same or different and selected from hydrogen, halogen, alkyl or alkoxy groups, wherein alkoxy groups may be branched or straight C_1-C_9-chains and the alkyl and alkoxy groups may comprise cyclic alkyl groups, for example cycloalkylalkyl.

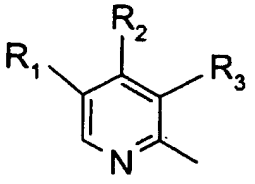
Preferably, the substituted sulfinyl heterocyclic compound containing an imidazole moiety prepared by the novel method is a magnesium salt of formula I'.



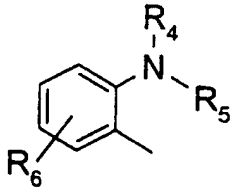
(I')

wherein

Ar is

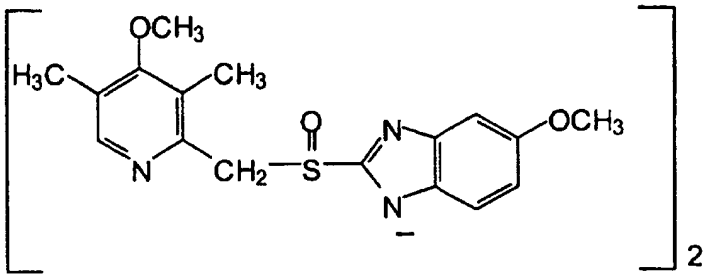


or

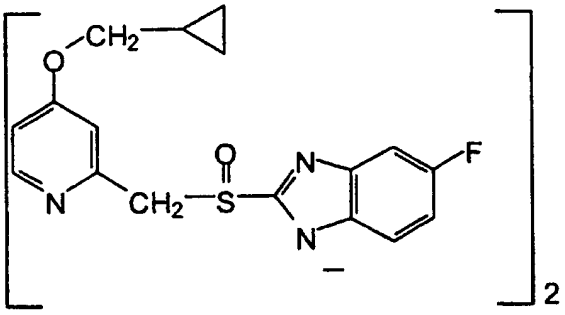


and R_1 - R_{11} are as defined above in connection with formula I.

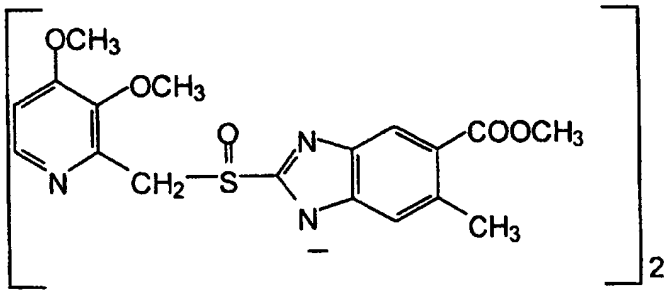
Most preferably the compounds prepared by the novel process are any of the formulas Ia to Ih.



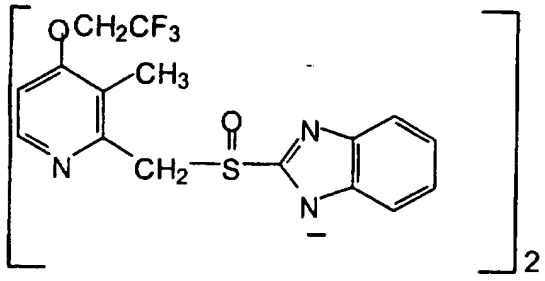
(Ia)



(Ib)

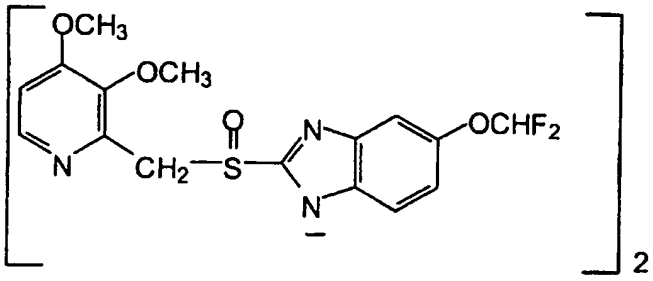


(Ic)



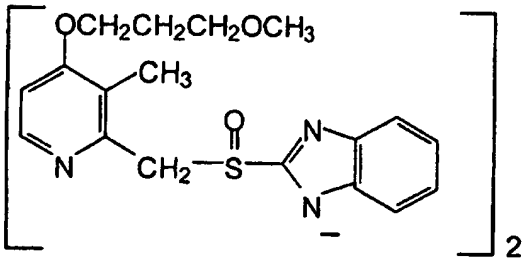
(Id)

Mg^{2+}



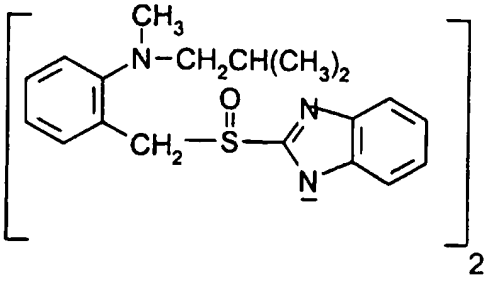
(Ie)

Mg^{2+}



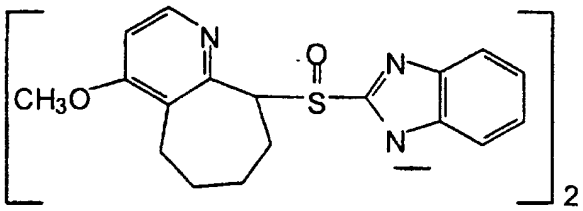
(If)

Mg^{2+}



(Ig)

Mg^{2+}



(Ih)

The substituted sulfinyl heterocycle of Formula I is mixed/reacted with a weak base and a magnesium source and optionally in the presence of an organic solvent. After the reaction is completed, the mixture is clarified, if needed. The product is preferably precipitated from the filtrate, optionally, by the addition of an appropriate solvent, for instance water or acetone, which facilities the precipitation of the product. As an additional benefit, when water is used, the solubility of the inorganic salts is enhanced resulting in less impurities in the form of inorganic salts in the obtained product. The obtained product may be further processed by recrystallization.

The novel process according to the present invention may be exemplified by the following reaction scheme showing a reaction between a substituted benzimidazole (HA) and a weak base (B) in the presence of a magnesium source (Mg_mX_n).

$$2HA + 2B \xrightarrow{\frac{1}{m}(Mg_mX_n)} MgA_2 + \frac{n}{m}\left[(HB)\frac{2m}{n} \bullet X\right]$$

In the above formula, wherein HA is a substituted benzimidazole, H denotes the most acidic proton in said compound, B is a weak base and X is a counterion to Mg^{2+} in the magnesium source (Mg_mX_n).

The base used in the reaction must not be toxic or it should only have a low toxicological effect. It shall preferably be a weak base to minimize precipitation of poorly soluble inorganic magnesium salts, such as magnesium hydroxide during the reaction sequence. Such precipitation of, for instance, magnesium hydroxide - is normally difficult to remove during the process and in the final product. With the expression weak base is meant a base with a pKa lower than alkoxides and hydroxides, but higher than the substituted sulfinyl

heterocycles of the present invention, preferably with a pKa from 7-12. More preferably the weak base is an organic amine or ammonia. With respect to environmental aspects the base shall preferably be one resulting in residues in the form of ammonium salts which easily can be isolated, for example by filtration or centrifugation, in order to minimize effluent of nitrogen based pollutants, such as ammonia.

The magnesium source may be an organic as well as an inorganic magnesium salt, such as magnesium acetate, magnesium nitrate, magnesium sulfate, magnesium carbonates and magnesium chloride, preferably magnesium sulfate.

If a solvent is used in the reaction, it is preferably one which can be used throughout the complete process. Such a solvent is preferably an alcohol, for instance methanol.

The process is not temperature sensitive and it may be carried out at ambient temperature. Of course the process temperature and time may be adjusted with respect to the quality and yield of the obtained product.

The new process according to the present invention may be exemplified in more general terms by the manufacture of omeprazole magnesium salt.

Omeprazole magnesium salt may be formed in accordance with the invention by treating a weight amount of omeprazole with weighed amounts of aqueous ammonia and magnesium sulfate in methanol.

The order of charging the different reactants is not critical for the produced product. A specific order may be preferred with respect to the equipment actually used in the factory.

The temperature may be -10°C to +50°C and preferably is between 0°C and ambient temperature. After termination of the reaction, the resulting inorganic magnesium salts are separated off in a suitable equipment, such as a centrifuge or a pressure filter.

The temperature of the clear solution is adjusted to -10°C to +40°C, preferably 10°C to 35°C. The solution may be seeded with omeprazole magnesium salt crystals and an amount of water is added to start the precipitation. The amount of water is not critical, but can be equal to or less than the volume of the solution; preferably the latter.

The formed crystalline product is separated from the mother liquid (filtrate), for instance by centrifugation or filtration. Other suitable procedures may be used to separate the product. The produced crystalline product is washed with aqueous methanol and dried under reduced pressure and heat.

The process according to the present invention is described in more detail by the following examples, which are not intended to limit the scope of the invention.

Examples

Example 1. Preparation of 5-methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

5-Methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (31.6 kg, 91.6 mol) together with aqueous NH_3 (7.4 kg, 107 mol) was added to methanol (212 l). To the obtained mixture $MgSO_4$ x 7 H_2O (17.6 kg, 69.9 mol) was added at ambient temperature. After the reaction was completed inorganic salts were removed by means of filtration. Water was added to the filtrate, the mixture was clarified and water (91 l) was added. The mixture was kept for stirring in order to crystallize the product. The obtained product was centrifuged and was washed with a mixture of MeOH/water. The product was dried at reduced pressure at 40 °C. Yield: 71%. (Mg content: found 3.47%, Theoretically calculated 3.41%)

The % crystallinity of the obtained product was measured with powder X-ray diffraction (XRD) as described bellow: A thin layer of the triturated sample was smeared onto a cut silicon single crystal zero background holder which was rotated during the measurement. Cu Kα radiation and constant or automatic antiscatter and divergence slits were used to obtain a diffractogram from 1 or 2° 2θ to at least 35°.

The % crystallinity was calculated with the formula

$$\% \text{ crystallinity} = 100*C / (A + C)$$

C= the area from the peaks in the diffractogram ("the crystalline area"),
A = the area between the peaks and the background ("the amorphous area").

Area calculations were performed between 4-33° 2θ. The lowest intensity value found in this interval was chosen as the constant background and subtracted from the area A. When constant slits were used the increased background at low angles due to the influence from the primary beam was also subtracted from the area A.

The crystallinity was measured to be 80 ± 5% (calculation interval 4 - 33°).

Example 2. Preparation of 5-methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

5-Methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (25 g, 72.4 mmol) together with isopropylamine (7.4 ml, 86.9 mmol) was added to methanol (100 ml). To the obtained mixture $MgSO_4$ x 7 H_2O (8.85 g, 35.9 mmol) was added at ambient temperature. After the reaction was completed inorganic salts were removed by means of filtration. Water was added to the filtrate, the mixture was clarified and water (100 ml) was added dropwise. The product was filtered off and was washed with a mixture of MeOH/water (50 ml, 1:1). The product was dried at reduced pressure overnight. Yield: 95%. (Mg-content: 3.41; calculated theoretically 3.41).

Example 3. Preparation of 5-methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

5-Methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (25 g, 72.4 mmol) together with isopropylamine (7.4 ml, 86.9 mmol) was added to methanol (100 ml). To the obtained mixture $Mg(O_{Ac})_2$ x 4 H_2O (9.34 g, 43.6 mmol) was added at ambient temperature. After the reaction was completed inorganic salts were removed by means of filtration. Water was added to the filtrate, the mixture was clarified and water (100 ml) was added dropwise. The obtained product was filtered off and was washed with a mixture of MeOH/water (50 ml, 1:1). The product was dried at reduced pressure overnight. Yield: 92%. (Mg content: 3.42; calculated theoretically: 3.41)

Example 4. Preparation of 5-methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

5-Methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (25 g, 72.4 mmol) together with isopropylamine (7.4 ml, 86.9 mmol) was added to methanol (100 ml). To the mixture $Mg(NO_3)_2$ x 6 H_2O (11.2 g, 43.7 mmol) was added at ambient temperature. After the reaction was completed inorganic salts were removed by means of filtration. Water was added to the filtrate, the mixture was filtered and the filter cake was washed with methanol (10 ml). Water (100 ml) was added dropwise to the

combined organic layers. The product was filtered off and was washed with a mixture of MeOH/water (50 ml, 1:1). The product was dried overnight. Yield: 89%. (Mg content: 3.39; calculated theoretically: 3.41))

Example 5. Preparation of 5-methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

5-Methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (1.0 g, 2.9 mmol) together with diethylamine (0.35 ml, 3.4 mmol) were added to methanol (9 ml). To the obtained mixture $MgCl_2$ (142 mg, 1.5 mmol) in methanol (2 ml) was added at ambient temperature. Water (6.5 ml) was added dropwise. The obtained product was filtered off and was washed with a mixture of MeOH/water (20 ml, 1:1). Yield: 76%. (Mg content: 3.38; calculated theoretically: 3.41)

Example 6: Preparation of (-)-5-fluoro-2[[(4-cyclopropylmethoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

(-)-5-Fluoro-2[[(4-cyclopropylmethoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (20 g, 57.9 mmol) together with NH_3 (7.5 ml, 100.2 mmol) was added to methanol (80 ml). To the mixture $MgSO_4 \times 7\ H_2O$ (11.4 g, 45.3 mmol) was added at ambient temperature. The mixture was clarified. Water (8 ml) was added dropwise during rapid stirring. Another portion of water (72 ml) was added dropwise for 75 minutes. The mixture was stirred for 50 minutes while the product precipitated. The product was filtered off and was washed with a mixture of MeOH/water (2 ml, 1:1). The product was dried at reduced pressure at 35 °C overnight. Yield: 61%. (Mg content: 3.40; calculated theoretically: 3.41)

Example 7: Preparation of 5-fluoro-2[[(4-cyclopropylmethoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

5-Fluoro-2[[(4-cyclopropylmethyoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (10 g, 28.9 mmol) together with isopropylamine (1.71 g, 28,9 mmol) was added to methanol (40 ml). To the obtained mixture $MgCl_2$ (1.35 g, 14 mmol) was added at ambient temperature. Excess of amine was evaporated off. The mixture was clarified and water (56.5 ml) was added dropwise. The mixture was cooled to 20 °C and the product was filtered off and was washed with a mixture of MeOH/water (20 ml, 3:1). The obtained product was dried at reduced pressure at 50 °C overnight. Yield: 86%. (Mg content: 3.42; calculated theoretically: 3.41)

Example 8: Preparation of 5-fluoro-2[[(4-cyclopropylmethoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt.

5-Fluoro-2[[(4-cyclopropylmethoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (690 g, 1.97 mol) together with aqueous NH_3 (140 ml, 2.17 mol) was added to methanol (2.4 l). To the obtained mixture $MgCl_2$ (105.2 g, 1.08 mol) in methanol (940 ml) was added. The mixture was clarified and water (350 ml) was added during rapid stirring. Another portion of water (3.15 l) was added and the mixture was stirred overnight. The product was filtered off and was washed with a mixture of MeOH/water (1 l, 4:1). Yield: 91%. (Mg content: 3.46; calculated theoretically: 3.41)

Example 9: Preparation of (-)-5-methoxy-2-[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt

(-)-5-Methoxy-2-[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (10.6 g, 29 mmol) together with aqueous ammonia (3.8 ml of 25%, 50 mmol) was added to methanol (40 ml). To the solution $MgSO_4 \times 7\ H_2O$ (5.7 g, 23 mmol) was added. After stirring for 10 minutes the mixture was filtered and the filtrate was diluted with methanol (60 ml). Acetone (150 ml) was added and the solution was seeded with crystals while stirring. After 14 hours the product was isolated by filtration and the

crystals were washed with methanol/acetone (50 ml). The product was dried over night. Yield: 41%. (Mg-content: found 3.33%, Calculated for $(C_{17}H_{18}N_3O_3S)_2Mg$ 3.41%).

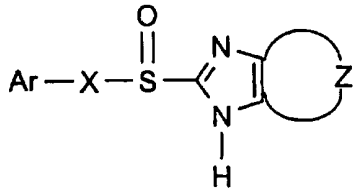
Example 10: Preparation of 5-difluoromethoxy-2-[[(3,4-dimethoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt

5-Difluoromethoxy-2-[[(3,4-dimethoxy-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole (11.1 g, 29 mmol) together with aqueous ammonia (3.8 ml of 25%, 50 mmol) was added to methanol (60 ml). To the solution $MgSO_4$ x 7 H_2O (5.7 g, 23 mmol) was added. After stirring for 3 minutes the mixture was filtered. Water (40 ml) was added dropwise to the filtrate while stirring. After 30 minutes the product was isolated by filtration and the crystals were washed with methanol/water (25 ml). The product was dried under reduced pressure. Yield: 67 %. (Mg-content: found 3.07%, Calculated for $(C_{16}H_{14}N_3O_4S)_2Mg$ 3.08%).

The best mode to practice the invention at present is by the process described in Example 1.

CLAIMS

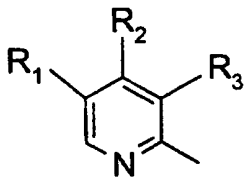
1. A process for the preparation of a magnesium salt of a substituted sulfinyl heterocyclic compound containing an imidazole moiety according to Formula I



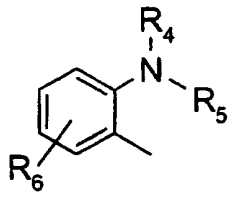
I

wherein

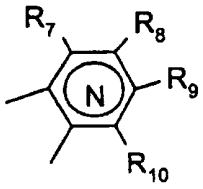
Ar is



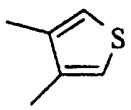
or



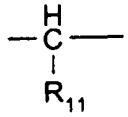
Z is



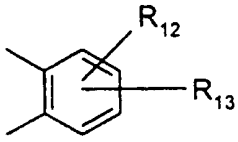
or



and X is



or



wherein

N inside the benzene ring of the benzimidazole moiety means that one of the carbon atoms substituted by R_7-R_{10} optionally may be exchanged for a nitrogen atom without any substituents;

R_1, R_2 and R_3 are the same or different and selected from hydrogen, alkyl, alkylthio, alkoxy optionally substituted by fluorine, alkoxyalkoxy, dialkylamino, piperidino, morpholino, halogen, phenylalkyl and phenylalkoxy, wherein alkyl and alkoxy groups may be branched or linear and may comprise cyclic alkyl groups such as cykloalkylalkoxi groups;

R_4 and R_5 are the same or different and selected from hydrogen, alkyl and aralkyl;

R_6 is hydrogen, halogen, trifluoromethyl, alkyl and alkoxy;

R_7 - R_{10} are the same or different and selected from hydrogen, alkyl, alkoxy, halogen, haloalkoxy, alkylcarbonyl, alkoxycarbonyl, oxazolyl, trifluoroalkyl, or adjacent groups R_7-R_{10} form ring structures which may be further substituted;

R_{11} is hydrogen or forms an alkylene chain together with R_3 and

R_{12} and R_{13} are the same or different and selected from hydrogen, halogen, alkyl or alkoxy groups, wherein alkoxy groups may be branched or straight C_1-C_9-chains and the alkyl and alkoxy groups may comprise cyclic alkyl groups, for example cycloalkylalkyl,

wherein the substituted sulfinyl heterocycle of Formula I is mixed together with a weak base and a magnesium source.

2. A process according to claim 1, wherein the weak base is selected from the group of organic amines and ammonia.

3. A process according to claim 1, wherein the weak base is ammonia.

4. A process according to claim 1, wherein the magnesium source is selected from the group of organic and inorganic magnesium salts.

5. A process according to claim 1, wherein the magnesium source is selected from the group of magnesium acetate, magnesium nitrate, magnesium sulfate, magnesium carbonates and magnesium chloride, preferably magnesium sulfate.

6. A process according to claim 1, wherein the reaction is carried out in the presence of a solvent.

7. A process according to claim 1, wherein the reaction is carried out in the presence of an aqueous organic solvent.

8. A process according to claim 1, wherein the weak base and magnesium source are selected to give an ammonium salt which can be removed by filtration during said process.

9. 5-Methoxy-2[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt prepared by a process according to any of claims 1 - 8.

10. (-)-5-Methoxy-2-[[(4-methoxy-3,5-dimethyl-2-pyridinyl)methyl]sulfinyl]-1H-benzimidazole, magnesium salt prepared by a process according to any of claims 1 - 8.

11. A pharmaceutical composition comprising a magnesium salt of a substituted sulfinyl heterocycle of formula I prepared by a process according to any of claims 1 - 8 as an active ingredient and a pharmaceutically acceptable carrier.

INTERNATIONAL SEARCH REPORT

International application No.
PCT/SE 97/00674

A. CLASSIFICATION OF SUBJECT MATTER

IPC6: C07D 401/12
According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC6: C07D

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

SE,DK,FI,NO classes as above

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

CAS-ONLINE

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	WO 9601623 A1 (ASTRA AKTIEBOLAG), 25 January 1996 (25.01.96)	1-8
X	--	9-11
A	WO 9501977 A1 (ASTRA AKTIEBOLAG), 19 January 1995 (19.01.95)	1-8
X	--	9,11

[X] Further documents are listed in the continuation of Box C. [X] See patent family annex.

* Special categories of cited documents:
"A" document defining the general state of the art which is not considered to be of particular relevance
"B" erlier document but published on or after the international filing date
"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
"O" document referring to an oral disclosure, use, exhibition or other means
"P" document published prior to the international filing date but later than the priority date claimed
"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
"X" document of particular relevance: the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
"Y" document of particular relevance: the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
"&" document member of the same patent family

Date of the actual completion of the international search	Date of mailing of the international search report
5 June 1997	07 -08- 1997
Name and mailing address of the ISA/ Swedish Patent Office Box 5055, S-102 42 STOCKHOLM Facsimile No. +46 8 666 02 86	Authorized officer Göran Karlsson Telephone No. +46 8 782 25 00

Form PCT/ISA/210 (second sheet) (July 1992)

INTERNATIONAL SEARCH REPORT

International application No.
PCT/SE 97/00674

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	Chemical Abstracts, Volume 108, No 9, 29 February 1988 (29.02.88), (Columbus, Ohio, USA), page 683, THE ABSTRACT No 75401p, JP, 62192365 A,, (Susumu et al) 22 August 1987 (22.08.87)	1-8
X	--	11

INTERNATIONAL SEARCH REPORT

Information on patent family members

20/05/97

International application No.
PCT/SE 97/00674

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
WO 9601623 A1	25/01/96	AU 2993795 A	09/02/96
		CA 2170647 A	25/01/96
		CN 1134666 A	30/10/96
		CZ 9600732 A	17/07/96
		EP 0723436 A	31/07/96
		FI 961057 A	29/03/96
		HU 9600573 D	00/00/00
		IL 114450 D	00/00/00
		JP 9502739 T	18/03/97
		NO 960950 A	07/03/96
		PL 313387 A	24/06/96
		SE 9402433 D	00/00/00
		ZA 9505548 A	08/01/96
		SE 9402432 D	00/00/00
WO 9501977 A1	19/01/95	AU 7198194 A	06/02/95
		BR 9406940 A	10/09/96
		CA 2166794 C	04/03/97
		CN 1126993 A	17/07/96
		CZ 9600069 A	15/05/96
		EP 0707580 A	24/04/96
		FI 960101 A	09/01/96
		HR 940385 A	28/02/97
		HU 9503873 D	00/00/00
		IL 110190 D	00/00/00
		JP 8512315 T	24/12/96
		NO 960068 A	05/01/96
		PL 312440 A	29/04/96
		SK 2296 A	01/10/96
		ZA 9404933 A	20/02/95
JP 62192365 A	22/08/87	NONE	